CHEM 112
PRINCIPLES OF CHEMISTRY
Lecture
Spring 2017

COURSE DESCRIPTION
An introductory course in chemistry emphasizing theoretical aspects and designed primarily for students who intend to take one or more additional courses in chemistry. The major topics covered are elementary classical thermodynamics (entropy and free energy), homogeneous and heterogeneous equilibria, properties of gases, liquids, and solids, theories of solutions and solubility, electrochemistry, acid-base theory and applications, and chemical kinetics.

Prerequisites: CHEM 111/111L; MATH 111 or equivalent,
Co-requisite: CHEM 112L

TEXTBOOK
Chemistry: Atoms First, 2nd Edition, by Burdge and Overby

COURSE OUTLINE
1. Ch 13  Physical Properties of Solutions; Colligative Properties
2. Ch 14  Entropy and Free Energy
3. Ch 15  Chemical Equilibrium
4. Ch 16  Acids, Bases and Salts
5. Ch 17  Acid-Base and Solubility Equilibria, Buffers
6. Ch 18  Electrochemistry
7. Ch 19  Chemical Kinetics and Catalysis

GENERAL INFORMATION
Meeting Times:
Lecture: Tuesday and Thursday, 1:40–2:55, SSMB 129

Instructor:
Dr. Richard Salinaro
Office: SSMB 326
Phone: 843-953-5587 (This is the Chemistry Department Office #, they will forward messages to Dr. Salinaro)
E-mail: salinaror@cofc.edu

Office Hours:
Dr. Salinaro will hold office hours on Tuesday and Thursday 3:15- 5PM. Also by appointment.

Messages:
Students may address questions and requests for appointments by contacting Dr. Salinaro via email (salinaror@cofc.edu). Emails will be responded to on or before the next business day.
Supplemental Instruction:
Supplemental instruction will be provided by Savannah Reaves (email: reavessm@g.cofc.edu). More information on the supplemental instruction program is available at http://csl.cofc.edu/supplemental-instruction.

Course Webpages:

OAKS: Important information regarding this course will be available on the OAKS webpage. This information includes all announcements and postings, lecture schedule and suggested reading assignments, select lecture notes, exam answer keys, handouts.

McGraw-Hill Connect: LearnSmart study modules and problem sets are available via the course’s Connect webpage. Homework will not be graded, it is your opportunity to hone your problem solving skills. However, Dr. Salinaro recommends you apply yourself here because this is your opportunity to practice. Students are required to purchase an access code for the companion textbook website, McGraw-Hill Connect Chemistry. An access code is bundled with the textbook and also can be purchased separately. Students can register and access section assignments at http://connect.mheducation.com/class/r-salinaro-tuesday-thursday-spring-2017

Attendance Policy and Classroom Conduct:
Students are expected to attend all classes, attendance will be taken randomly. Students are responsible for all information presented in class whether they are present or not. Students should obtain notes from a classmate and read the associated material in the text BEFORE they request help from the instructor about material missed.

Please note that an Absence Memorandum from the Office of Undergraduate Studies only verifies your documentation for missing a class. It does not entitle you to make up or be excused from any work, assignment or test.

In order to foster a cordial and secure learning environment, please be respectful of your instructor and your classmates:

- Arrive on time
- No cell phone use in class, Set cell phones on mute or vibrate before coming to lecture. Students who have difficulty with the cell phone policy will be asked to place their phones on their desks in red plastic bags.
- Do not obstruct or disrupt the teaching and learning processes by carrying on conversations on your cell phone or with other students in the class, sending text messages, or surfing the web on your laptop.
- Do not verbally abuse, threaten, intimidate, or ridicule your instructor or classmates.
- No eating in class.

If you fail to comply with these simple requests you will be asked to leave the class and if the problems persist you will be referred to the Dean of Students for disciplinary action.
Exams:

Four exams will be given. Ample notice will be given prior to an exam. Absences from any exam must be arranged in advance. No make-up exams will be given. Students should contact Dr. Salinaro as soon as possible regarding scheduling conflicts.

Answer keys will be posted on the course webpage.

Quizzes:

In-class quizzes will be administered throughout the semester. There will be two types of quizzes: Type 1 - a one question, open notes, unannounced quiz given at the start of class, or Type 2 - a prescheduled quiz consisting of two or three questions. You will be given ample notice that a Type 2 quiz is being given. Consult your instructor if there is a scheduling conflict. Absences from a Type 2 quiz must be arranged in advance. No make-up quizzes will be given other than for acceptable reasons such as illness (the student must provide documentation, e.g., an absence memo), attending a conference, or circumstances beyond the student's control.

Final Exam:

A 110 minute comprehensive and standardized ACS final exam is scheduled for 12-3PM Tuesday, May 2. The multiple-choice final exam is cumulative over the material covered during the entire semester.

Requests for an alternate final exam time must be processed through the Office of Undergraduate Studies no later than 5 p.m. on the last day of class. Failure to take the final exam will result in a grade of "X" which turns to an "F" after 48 hours. Undergraduate students should be aware that excuses for missing final examinations may be obtained from the Office of Undergraduate Studies.

The acceptable reasons for missing final examinations are illness of the student (the student must provide documentation, e.g., an absence memo) or circumstances beyond the student's control. These reasons must be properly documented. See the section entitled "Final Examinations" in the Undergraduate Catalog for more information. Examinations must be taken at the time scheduled except when [a] two or more exams are scheduled simultaneously, or [b] the student has three examinations within a 24-hour period.

Electronics Device Policy:

No electronic devices except for calculators are allowed during exams. The use of any wireless communication devices, iPhones, iWatches, etc., during a test or the final exam is prohibited and will be considered to be a violation of the Honor Code.

“Make-up” Policy and Regrades:

As mentioned above, there are no make-up-tests for missed exams under any circumstances. If you have an emergency that is documented with a note from the Dean of Undergraduate Studies or a note from a medical doctor, you may be excused from an exam. Contact me as soon as possible.

Students may return exams they believe to have significant grading errors for reconsideration within one week of receipt of the graded exam. A significant error will
constitute a > 3 pt mistake on grading of an individual question. Students must submit clear and succinct explanations of the grading error(s) in question along with the exam to be regarded. The explanation should establish that the answer key is incorrect or incomplete, that the answer given by the student is an equivalent or equally valid solution to that given on the key, or that the student gave the same answer as the key but it was not recognized as such. No markings or other alterations should be made on the exam itself. To ensure fair and equal treatment to all students, all changes in exam scores will be made only through this formal re-grade process. Dr. Salinaro will not discuss exam-score changes nor make exam-score changes in face-to-face meetings with students.

HONOR CODE

Student conduct is expected to conform to the standards of the College of Charleston Student Honor Code Policy. In addition, students in this course are also expected to be aware and to conform to the standards of the Department of Chemistry & Biochemistry Policy on Scientific Integrity.

Students at the College of Charleston are bound by honor and by their enrollment at the College to abide by the Honor and Conduct codes and to report violations. Faculty and staff members are equally required to report violations of the Honor Code or Code of Conduct. Students violating the Honor Code will be remanded to the college’s Honor Board.

STUDENT DISABILITY/ACCESS STATEMENT

This College abides by Section 504 of the Rehabilitation Act of 1973 and the Americans with Disabilities Act and will make reasonable accommodations for persons with documented disabilities. Students should apply for services at the Center for Disability Services/SNAP located on the first floor of the Lightsey Center, Suite 104, (843) 953-1431. If you have a documented disability that may have some impact on your work in this class and for which you may require accommodations, you are responsible for notifying me as soon as possible and for contacting me one week before accommodation is needed.
LEARNING OBJECTIVES

General chemistry provides you with an opportunity to do lots and lots of quantitative (numerical) and qualitative (conceptual) reasoning. General chemistry provides an introduction to the same material that is covered in the upper level physical chemistry courses, i.e., CHEM 341 and CHEM 342. This course also provides an important foundation for CHEM 231 and CHEM 232. At the end of CHEM 112 you should be able to use the tools of basic mathematics and physics to solve problems in chemistry and biochemistry. If someone poses a question about the physical basis of some chemical phenomenon, you should be able to apply your knowledge to suggest the appropriate theory or model to apply, be able to do the calculations necessary to apply the model and explain what you have done clearly and coherently so the person who asked the question has confidence that you know what it is you are doing. It would also be nice if, along the way, you gain some appreciation for the underlying beauty of the physical world. However, progress on this goal is hard for me to assess, so I’ll leave it up you to decide how you are coming along.

At the end of this syllabus is a detailed list of subjects and knowledge requirements that you will encounter throughout the semester. Listed below are the general (and rather broad) learning objectives for CHEM 112:

1. Demonstrate competency with all of the learning objectives stated for Chem 111 and Math 111.
2. Apply common mathematical techniques to describe the kinetic and thermodynamic processes related to chemical equilibria.

GENERAL EDUCATION LEARNING OUTCOMES

1. Students apply physical/natural principles to analyze and solve problems.
2. Students explain how science impacts society.

The General Education Learning Outcomes will be assessed in the Lab Co-requisite course, Chem 112L
SEMESTER GRADES

Semester grades will be calculated by weighted average using the following criteria:

<table>
<thead>
<tr>
<th>Grade Percentage</th>
<th>Description</th>
</tr>
</thead>
<tbody>
<tr>
<td>LearnSmart Prep Assignment</td>
<td>5%</td>
</tr>
<tr>
<td>LearnSmart Homework</td>
<td>10%</td>
</tr>
<tr>
<td>Quizzes</td>
<td>15%</td>
</tr>
<tr>
<td>Exams</td>
<td>50%</td>
</tr>
<tr>
<td>Final Exam</td>
<td>20%</td>
</tr>
</tbody>
</table>

Your final average is a weighted average score based on the percentages shown in the above table. Letter grades will be assigned based on the straight grading scale shown in the table below.

<table>
<thead>
<tr>
<th>Score/%</th>
<th>Grade</th>
</tr>
</thead>
<tbody>
<tr>
<td>93-100</td>
<td>A</td>
</tr>
<tr>
<td>90-92</td>
<td>A-</td>
</tr>
<tr>
<td>87-89</td>
<td>B+</td>
</tr>
<tr>
<td>83-86</td>
<td>B</td>
</tr>
<tr>
<td>80-82</td>
<td>B-</td>
</tr>
<tr>
<td>77-79</td>
<td>C+</td>
</tr>
<tr>
<td>73-76</td>
<td>C</td>
</tr>
<tr>
<td>70-72</td>
<td>C-</td>
</tr>
<tr>
<td>67-69</td>
<td>D+</td>
</tr>
<tr>
<td>64-66</td>
<td>D</td>
</tr>
<tr>
<td>61-63</td>
<td>D-</td>
</tr>
<tr>
<td>&lt; 61</td>
<td>F</td>
</tr>
</tbody>
</table>

The instructor reserves the right to increase a student’s grade if the instructor feels that it is warranted. Periodically you will be made aware of your class standing so that you can assess your progress and to help you avoid any surprises at the end of the semester. However, it is suggested you keep a record of your performance on tests and quizzes to stay current with your standing in class.